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Modern Chemistry Chapter 6 2

Modern Chemistry Chapter 6 Section 2. STUDY. PLAY. single bond. is a covalent bond in which one pair of electrons is shared between two atoms. multiple bond. double and triple bonds are referred. molecule. a neutral group of atoms that are held together by covalent bonds. molecular compound.

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CHAPTER 6 REVIEW Chemical Bonding
SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the.

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Modern Chemistry 1 Chemical Bonding

CHAPTER 6 Chemical Bonding SECTION

1 Introduction to Chemical Bonding

OBJECTIVES 1. Define Chemical bond. 2.

Explain why most atoms form chemical

bonds. 3. Describe ionic and covalent

bonding.. 4.

CHAPTER 6 Chemical Bonding

Modern Chemistry Chapter 6. STUDY.

PLAY. Chemical Bond. Forces that hold

groups of atoms together and make

them function as a unit. Ionic Bonding.

Chemical bonding that results from the

electrical attraction between cations and

anions. 1.7 and 1.7+ (Atoms completely

give up electrons to other atoms)

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CHAPTER 6 REVIEW Chemical Bonding

SECTION 2 SHORT ANSWER Answer the

following questions in the space

provided. 1. Use the concept of potential

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energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

6 Chemical Bonding

Section 6.2 Classification of the Elements 1. h 2. f 3. n 4. a 5. k 6. d 7. l 8. o 9. i 10. g 11. b 12. m 13. c 14. j 15. e 16. Sodium and potassium have the same number of valence electrons. 17. The energy level indicates the period. For example, lithium's valence electron is in the second energy level and lithium is found in period 2. 18. four 19. groups 1 and 2 20.

Ch 6 Study Guide answers

The electronegativities are 2.5 for S, 2.1 for H, 0.7 for Cs, and 3.0 for Cl. In each pair, the atom with the greater electronegativity will be the more negative atom. Bonded atoms
Electronegativity difference Bond type
More-negative atom S and H $2.5 - 2.1 = 0.4$ polar-covalent S S and Cs $2.5 - 0.7$

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= 1.8 ionic S

CHAPTER 6 Chemical Bonding

Play this game to review Chemistry. An ionic compound is NOT represented by a molecular formula because an ionic compound: ... An ionic compound is NOT represented by a molecular formula because an ionic compound: Holt Modern Chemistry Chapter 6.3 Section Quiz DRAFT. 10th - 12th grade. 0 times. Chemistry. 0% average accuracy. 18 minutes ago ...

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HOMEWORK 6-3 (pp. 167-169)

VOCABULARY Write true or false for each
statement. 1. The bond length is the
distance between two bonded atoms at
their maximum potential energy. ____ 2.
The energy required to form a chemical
bond between two neutral atoms is
called bond energy. ____ 3.

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HOMEWORK 6-1

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