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Stoichiometry Guide Answers

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weight of an element to convert from the mass of the compound we have to the molar amount of it. A “mole” is a specific number of something —  $6.022 \times 10^{23}$ .

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SOLUTION: You are asked to calculate the amount (in moles) of oxygen atoms in a given sample. You are given the mass of the oxygen sample. Use the molar mass of oxygen ( $1 \text{ mol O} = 16.00 \text{ g}$ ) to create a conversion factor that

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converts mass (in grams) to amount (in mol) of oxygen.  $124 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g}} = 7.75 \text{ mol O}$ .

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who as well as don't subsequently  
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Chemistry Chemical compounds with  
integral atomic ratios, like nitrous oxide,

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are described as stoichiometric compounds, and they permit many simple calculations. The common oxide of aluminum provides a second example, but this time, begin with the mass percent and deduce the atomic ratio.

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are chemically engineered, or produced through chemical reactions. Chemically engineered commodities all rely on stoichiometry for their production.

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Stoichiometry. Limiting reactant. excess

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reactant. Stoichiometric coefficients. The method used to relate the quantities of different substances.... The reactant that runs out first in a chemical equation, thus.... The reactant that is not completely used up in a chemical reaction....

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Stoichiometry is a section of chemistry that involves using relationships between reactants and/or products in a chemical reaction to determine desired quantitative data. In Greek, stoikhein means element and metron means measure, so stoichiometry literally translated means the measure of

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elements.

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4, 72.08 10. methanol and oxygen gas  
11. carbon dioxide and water 12. 160.10  
g 13. 160.10 g 14. They are equal. 15. A  
mole ratio is a ratio between ...

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## **Website**

Section 10.1 shows the general equation  
stoichiometry steps as measurable  
property 1  $\square$  moles 1  $\square$  moles 2  $\square$   
measurable property 2 When the  
reactants and products of a reaction are  
pure solids and pure liquids, mass is the  
conveniently measurable property, but  
many chemical changes take place in

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either the gas phase or in solution.

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