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mole ratio. excess reactant. limiting
reactant. The study of quantitative

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relationships between the amounts of....
In a balanced equation, the ratio
between the numbers of moles.... A
reactant that remains after a chemical
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barometer. standard atmospheric

pressure. manometer. daltons law of

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partial pressures. a device that
measures atmospheric pressure. It is
filled with.... $1 \text{ atm} = 760 \text{ mm Hg}$...
 $\text{pascal} = 1 \text{ N/m}^2$... $1 \text{ atm} = 101.3$...
 $\text{kpa} = 101,325 \text{ Pa}$. measures gases in a
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barometer. standard atmospheric
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between the amounts of reactants used and the amounts of products formed in a chemical reaction. Mass of Reactants = Mass of Products. ... Chemistry Matter and Change: Chapter 12 - States of Matter. 16 terms.

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Molality moles of solute L of solution
moles of solute kg of solvent Percent
mass grams of solute $\times 100$ grams of
solution Mole Fraction mol solute total
moles Density of Solution grams solution
mL solution X Solution Formation
Electrolytic ions present in solution

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Hydrated ions. interactions keep the ions from recrystallizing Not everything that dissolves in water is an electrolytic solution. o Sugar dissolves.

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1. Choose the element in the set that you would expect to be most metallic. S,

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Cl, Kr, Xe, Po 2. How many core electrons are there in an Sb ($Z = 51$) atom?

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