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Ne = 90%;  $^{22}\text{Ne}$  = 10% Q5. The A r of lithium is 6.92. The most common isotope is  $^7\text{Li}$  with an abundance of 92%. Determine the mass number of the other isotope.

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the 1. 2. 4. 5. 6. 8. 9. The 3 particles of the atom are: protons ( $e^-$ ) (o Their respective charges are: The number of protons in ...

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$$\text{atomic mass} = (78.6 \times 24) + (10.1 \times 25) \\ + (11.3 \times 26) \dots$$

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For hydrogen, the single electron is placed in the 1s orbital, which is the orbital lowest in energy (Figure 1.2.5), and the electron configuration is written

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as  $1s^1$  and read as "one-s-one." A neutral helium atom, with an atomic number of 2 ( $Z = 2$ ), has two electrons.

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number, mass number, and isotopes.

Practice: Isotope composition: Counting protons, electrons, and neutrons.

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metallic plate is the: ... Isotopes (such as hydrogen-1, hydrogen-2, and hydrogen-3) are atoms of the same element that differ in: the number of protons in the nucleus. the number of electrons. the number of neutrons in the nucleus.

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