Chapter 9 Review Stoichiometry Section 3 Answers Modern Chemistry

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6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g) → 2NH 3(g) N

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following

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Review ...

Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through “File info” CorrectionKey=NL-A

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

Stoichiometry Worksheet Answers Chapter 9

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