

Chapter 8 Review Chemical Equations And Reactions Answers

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CHAPTER 8 REVIEW Chemical Equations and Reactions SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg. 2. Consider the metals iron and silver, both listed in Table 3on page 286 of the text. Which one

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Chemistry Chapter 8 Review: Chemical Equations ____1. In a chemical equation, the __1__ are found on the left side of the arrow, ____2. while, the __2__ are found ...

Chemistry Chapter 8 Review: Chemical Equations

Modern Chemistry 1 Chemical Equations and Reactions CHAPTER 8 REVIEW Chemical Equations and Reactions Teacher Notes and Answers Chapter 8 SECTION 1 SHORT ANSWER 1. a. d b. a c. b d. f e. e f. c 2. 8,4,9 3. a. 12 atoms b. 16 atoms c. 51 atoms d. 3 1024 atoms 4. $2\text{Al}(s) + 3\text{CuF}_2(aq) \rightarrow 2\text{AlF}_3(aq) + 3\text{Cu}(s)$ 5. $\text{NaCl}(aq) + \text{AgNO}_3(aq) \rightarrow \text{AgCl}(s) + \text{NaNO}_3(aq)$ 6. a.

CHAPTER 8 REVIEW Chemical Equations and Reactions

CHAPTER 8 REVIEW Chemical Equations and Reactions MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. b A balanced chemical equation represents all the following except (a) experimentally established facts. (b) the mechanism by which reactants combine to form products.

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which reactants combine to form products.

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Chemical Equations and Reactions The evolution of energy as light and heat is an indication that a chemical reaction is taking place. CHAPTER 8 Fireworks. CHEMICAL EQUATIONS AND REACTIONS 261 SECTION 1 OBJECTIVES List three observations that suggest that a chemical reaction has taken place.

CHAPTER 8 Chemical Equations and Reactions

Balanced chemical equations show that the elements, atoms and mass on the left equals the elements, atoms and mass on the right: the matter that goes in must come out. Example: The following chemical equation shows that ammonium dichromate breaks down into nitrogen, chromium(III) oxide, and water.

St. Charles Parish Public School System

Katalynn Sutton Chapter 8 Review Chemical Equations 8.1 A chemical equation is a shorthand expression for a chemical change or reaction which uses chemical symbols and formulas of the reactant and products and other symbolic terms to represent a chemical reaction.

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CHAPTER 8 REVIEW Chemical Equations and Reactions SECTION 3 SHORT ANSWER Answer the following questions in the space provided 1 List four metals that will not replace hydrogen in an acid Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg 2 Consider the metals iron and silver, both listed in Table 3on page 286 of the text Which one Math Tutor Balancing ...

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CHAPTER 8 REVIEW Chemical Equations and Reactions SECTION 2 SHORT ANSWER Answer the following questions in the space provided 1 Match the equation type on the left to its representation on the right c synthesis (a) $AX + BY \rightarrow AY + BX$ d decomposition (b) $A + BX \rightarrow AX + B$ b single-displacement (c) $A + B \rightarrow AX$ a double-displacement (d) $AX \rightarrow A + X$

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Name Date CHAPTER 8 REVIEW Class S E-c7b IV hemical Equations and Reactions MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. b d A balanced chemical equation represents all the following except (a) experimentally established facts. (b) the mechanism by which reactants combine to form products.

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8. 9. 10. The actual knowledge about what products are produced In a chemical reaction IS

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obtained by (a) inspecting the chemical equation; (b) balancing the chemical equation; (c) laboratory analysis; (d) Balancing a word equation. Once the correct formula for a reactant in an equation has been written, the

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