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compounds.
nomenclature.

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oxidation states. an ion formed from a single atom. (remember to change nonmetal.... compounds composed of two different elements; (compounds of 2.... naming system. the condition of an atom expressed by the number of electrons....

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Question 1. A liquid is
in equilibrium with its

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vapours in a sealed container at a fixed temperature. The volume of the container is suddenly increased, (i) What is the initial effect of the change on the vapour pressure ?

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7.1 Introduction: Recall from Chapter 1 that solutions are defined as homogeneous mixtures that are mixed so thoroughly that neither component can be observed independently of the other. Solutions are all around us. Air, for example, is a solution. If you live near a lake, a river, or an ocean,

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that body of water is not pure H_2O but most probably a solution.

CH104: Chapter 7 - Solutions - Chemistry

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of all the
electrons, then show
the electron
configuration of the

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electron using the
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technique for each of
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Spectrum. Energy
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Chapter 7: P Block Elements is one of the lengthiest chapters of class 12 chemistry syllabus. There are a variety of concepts and topics in this chapter. Thus to tackle this problem we need to start studying in an organized way by understanding and relating the concepts.

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classified as organic
and inorganic
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their origin. Organic compounds are obtained from living things whereas [...]

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For example, the density of water at its critical point ($T = 374^{\circ}\text{C}$, $P = 217.7 \text{ atm}$) is 0.32 g/mL , about one-third that of liquid water at room temperature but much greater than that of water vapor under most conditions. The transition between a

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liquid/gas mixture and a supercritical phase is demonstrated for a sample of chlorine in Figure 7.6.1 ...

Chapter 7.6: Critical Temperature and Pressure - Chemistry

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