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Chapter 6 Chemical Bonding Section

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

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Bonding Section 2

ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

6 Chemical Bonding

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Chapter 6 Chemical

Bonding Bonding

between Electroneg.

More-neg- sulfur and

difference Bond type

ative atom. hydrogen

$2.5 - 2.1 = 0.4$ polar-

covalent sulfur cesium

$2.5 - 0.7 = 1.8$ ionic

sulfur chlorine $3.0 - 2.5$

$= 0.5$ polar-covalent

chlorine. Chemical

Bonding, continued.

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Chapter 6 Chemical Bonding Table of Contents

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Introduction to
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Terms chemical bond
nonpolar-covalent
bond ...

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Modern Chemistry 9
Chemical Bonding
CHAPTER 6 STUDY
GUIDE Chemical
Bonding SECTION 3
IONIC BONDING AND
IONIC COMPOUNDS
SHORT ANSWER

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Answer the following questions in the space provided. 1. _____ The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal.

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IONIC BONDING

COVALENT BONDING

Atoms A Atoms B Atom

C Atom D Electrons

transferred from atoms

A to atoms B Electron

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pair shared between
atom C and atom D +

+ Many atoms Two

atoms Atom C Atom D

Cation A Anion B + + +

+ + + +----- - SECTION

6.1 Nature favors

arrangements in which

potential energy is

minimized. For

example, a boulder is

less likely to balance

CHAPTER 6 Chemical Bonding

Chapter 6: Chemical

Bonding Section 1-

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Introduction to
Chemical Bonding
Objectives: define
chemical bond;
differentiate between
covalent and ionic
bonding; explain why
bonding occurs; use
the difference in
electronegativity to
determine whether a
bond is polar covalent,
nonpolar covalent or
ionic

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Chapter 6 – Chemical Bonds. Jennie L. Borders. Standards. SPS1. Students will investigate our current understanding of the atom. b. Compare and contrast ionic and covalent bonds in terms of electron movement. SPS2. Students will explore the nature of matter, its classification and its system for naming types of matter.

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**Chapter 6 - Chemical
Bonds**

Modern Chemistry 41

Chemical Bonding

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

_____ A chemical bond
between atoms results
from the attraction
between the valence
electrons and _____ of
different atoms. (a)
nuclei (c) isotopes (b)

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inner electrons (d)

Lewis structures 2.

CHAPTER 6 REVIEW Chemical Bonding

A hydrogen bond is a dipole - dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up

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or shared, the
hydrogen bond is a
weaker attraction.

Chapter 6 Review: Chemical Bonding Flashcards | Quizlet

A chemical bond
between atoms results
from the attraction
between electrons and.
A shared electron pair.
A covalent bond
consists of. Nonpolar
covalent. If two
covalently bonded
atoms are identical,

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the bond is identified as. Polar. A covalent bond in which there is an unequal attraction for the shared electrons is.

Chapter 6 Section 6-1 Review

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6 Chemical Bonding.

CHAPTER 6 REVIEW.

Chemical Bonding.

SECTION 2. SHORT

ANSWERS

Answer the following questions in

the space provided. 1.

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Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the electron-proton attraction is stronger than the electron-electron and proton-proton repulsions.

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Chapter 6 Notes -

Chemical Bonding

Chemical bond - A

mutual electrical

attraction between the

nuclei and valence

electrons of different

atoms that binds the

atoms together 6-1

Introduction to

Chemical Bonding I.

Types of Chemical

Bonding A. Ionic

Bonding 1.

Chapter 6 Notes -

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Covalent Answer

Key Section 2 Covalent

Answer Key malleable and ductile but ionic-crystalline compounds are not. The metallic bond is the same in all directions throughout the metallic structure allowing the atoms to slide past each other. This sliding is why metals are ductile and malleable.

Chapter 6 Chemical

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Bonding Section 2
**Bonding Section 2
Covalent Answer
Key**

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1.

- a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d)

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Lewis structures 2. b A covalent bond consists of (a) a shared electron.

Section 6 1 Introduction To Chemical Bonding Answer Key

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Bonding Modern
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Bonding CHAPTER 6
REVIEW Chemical
Bonding SECTION 4
SHORT ANSWER

Answer the following

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questions in the space provided. 1. In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in ...

Chapter 6 Review Chemical Bonding Answer Key

As you read in Section 6-1, nature favors chemical bonding

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Bonding Section 2

because most atoms are at lower potential energy when bonded to other atoms than they are at as independent particles. In the case of covalent bond formation, this idea is illustrated by a simple example, the formation of a hydrogen-hydrogen bond.

CHAPTER 6 Chemical Bonding

Chapter 6 Chemical

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Bonds Section 6.2

Covalent Bonding

(pages 165–169) This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar.

Chapter 6 Chemical Bonds Section 6.2 Covalent Bonding

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Chapter Exam

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Instructions. Choose your answers to the questions and click 'Next' to see the next set of questions.

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