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SECTION 3 SHORT ANSWER Answer the following questions in the space provided.

1. c When an electron is added to a neutral atom, energy is (a) always absorbed. (c) either absorbed or released. (b) always released. (d) neither absorbed nor released.

2. d The energy required to remove an electron from a

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neutral atom is the atom's

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1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design.

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SHORT ANSWER Answer the following questions in the space provided. 1.

Consider the neutral atom with 53 protons and 74 neutrons to answer the following questions. 53. a. What is its

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atomic number? 127. b. What is its mass number? atomic number. c. Is the element's position in a modern periodic table ...

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SECTION 5-3. SHORT ANSWERAnswer

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the following questions in the space provided. 1. When an electron is added to a neutral atom, energy is . (a)always absorbed(c)either absorbed or released. (b)always released(d)burned away. 2.

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Science Chapter 5 The Periodic Table.

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The atomic mass of an element is. Which statement is true about the metal.... At room temperature none of the metals.... Which general statement does NOT apply.... a weighted average of the masses of an element's isotopes. Silicon is a better conductor of electric current than sulfur....

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absorbed (c) either absorbed or released  
(b) always released (d) burned away 2. d  
The energy required to remove an  
electron from an atom is the atom's .



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absorbed (c) either absorbed or released  
(b) always released (d) burned away 2

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The energy required to remove an electron from an atom is the

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## **Chapter 5 2 Review The Periodic**

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## **Law Answers**

SECTION 1 History of the Periodic Table  
KEY TERMS • The periodic law states that the physical and chemical properties of the elements are periodic functions of their atomic numbers. • The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar

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properties fall in the same column.

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